**The World of Stoichiometry!**

Use the following unbalanced equation to answer the questions below:

\_\_\_\_ Ca(OH)2 + \_\_\_\_ HBr → \_\_\_\_ H2O + \_\_\_\_ CaBr2

1. How many grams of calcium bromide can I make from 53 grams of calcium hydroxide and an excess of hydrobromic acid?
2. How many grams of HBr will be needed to make 118 grams of calcium bromide?
3. If I do this reaction with 35 grams of calcium hydroxide and 45 grams of hydrobromic acid, how many grams of calcium bromide will I make? What is the limiting reagent?